Supporting Information

Effects of Spiro-Cyclohexane Substitution of Nitroxyl Biradicals on Dynamic Nuclear Polarization

Nargiz B. Asanbaeva^a, Larisa Yu. Gurskaya^a, Yuliya F. Polienko^a, Tatyana V. Rybalova^a, Maxim S. Kazantsev^a, Alexey A. Dmitriev^b, Nina P. Gritsan^b, Nadia Haro-Mares^c, Torsten Gutmann^c, Gerd Buntkowsky^c, Evgeny V. Tretyakov^d, Elena G. Bagryanskaya^{a*}

^aN.N. Vorozhtsov Institute of Organic Chemistry, 9 Ac. Lavrentiev Avenue, Novosibirsk 630090, Russian Federation

^bV.V. Voevodsky Institute of Chemical Kinetics and Combustion, Siberian Branch of the Russian Academy of Sciences, 3 Institutskaya Str, 630090 Novosibirsk, Russian Federation

^cTU Darmstadt; Eduard-Zintl-Institute for Inorganic and Physical Chemistry, Alarich-Weiss-Straße 8, 64287 Darmstadt, Germany

^dN.D. Zelinsky Institute of Organic Chemistry, Russian Academy of Sciences, 47 Leninsky prosp., 119991 Moscow, Russian Federation

X-ray crystallography

XRD data were obtained at room temperature on a Bruker Kappa Apex II CCD diffractometer with Mo K α radiation (λ = 0.71073 Å) and a graphite monochromator. Absorption corrections were applied empirically using *SADABS* programs [1]. The structures were solved by direct methods using the *SHELX-97* [2] programs set and refined by full-matrix least-squares method against all F^2 in anisotropic (beside the atoms H) approximation using the *SHELXL2014*/7 programs set [3]. The H atoms positions were calculated with the riding model. The asymmetric unit contains also molecule of chloroform. The obtained crystal structure geometry and short contacts between non-bonded atoms were analyzed using the PLATON program [4].

Crystal solvate of diradical **4** is triclinic, space group P-1, a 12.4262(5), b 13.3902(5), c 13.7759(6) Å, α 67.121(2), β 67.676(2), γ 77.066(2)°, V 1945.7(1) Å³, Z 2, $C_{41}H_{51}C_{13}N_6O_2Fe$, Dc 1.333 g/cm³, μ 0.638 mm⁻¹, F(000) 864, crystal size 0.20 × 0.06 × 0.05 mm³, independent reflections 8868, R 0.0973,R2 0.2002 for, S 1.028 for all reflections (R 0.0657, wR2 0.1767 for 6296 I>2 σ), largest difference peak and hole 1.769 and -1.412 e.Å⁻³.

CCDC 2168987 contains the supplementary crystallographic data for this paper. These data can be obtained free of charge *via* http://www.ccdc.cam.ac.uk/cgi-bin/catreq.cgi, or from the Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: (+44) 1223 336 033; or e-mail: deposit@ccdc.cam.ac.uk

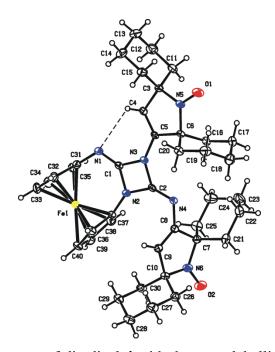


Figure S1. Molecular structure of diradical **4** with the atom-labelling scheme. Displacement ellipsoids are drawn at the 30% probability level.

Intermolecular short contact Cl3...O2 equaling to 3.067(3) Å is observed in solvate of **4** (sum of VDW radii [5] is 3.34 Å) (Figure S2). Solvates are bound through H-bond C1S-H...N1 (Table 3) between chloroform and 2,4-diimino-1,3-diazetidine moiety into chains directed along the (b-c). The chains, in turn, are formally bound via C40-H...O1 h-bond in (b, c) plain and then via C32-H...O1 and C38-H...O2 h-bonds along axis a into 3D-structure.

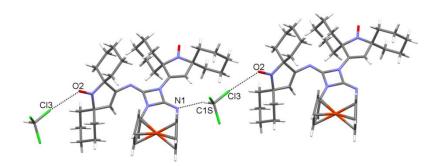


Figure S2. Fragment of molecular chain directed along (*b-c*).

2.61

H...O, Å C...O, Å C-H...O, °

C1S-H...N1 2.59 3.536(6) 161

C32-H...O1 2.59 3.485(6) 163

C40-H...O1 2.51 3.274(5) 140

3.403(5)

144

Table S1. Intermolecular hydrogen bonds in crystals of diradical 4 solvate.

References

C38-H...O2

- 1. Sheldrick, G.M. SADABS, Program for Empirical Absorption Correction of Area Detector Data; 1996;
- 2. Sheldrick, G.M. A Short History of {\it SHELX}. *Acta Crystallogr. Sect. A* **2008**, *64*, 112–122, doi:10.1107/S0108767307043930.
- 3. Sheldrick, G.M. Crystal Structure Refinement with SHELXL. *Acta Crystallogr. Sect. C Struct. Chem.* **2015**, *71*, 3–8, doi:10.1107/S2053229614024218.
- 4. Spek, A. Structure Validation in Chemical Crystallography. *Acta Crystallogr. D. Biol. Crystallogr.* **2009**, *65*, 148–155, doi:10.1107/S090744490804362X.
- 5. Rowland, R.S.; Taylor, R. Intermolecular Nonbonded Contact Distances in Organic Crystal Structures: Comparison with Distances Expected from van Der Waals Radii. *J. Phys. Chem.* **1996**, *100*, 7384–7391, doi:10.1021/jp953141+.